**Warm-Up 6.2**

**“Boiling Balloon”**

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| **Topic** | Gases and their Properties |
| **Benchmark**[**SC.CH.3.5**](http://165.248.72.55/hcpsv3/imr/report_by_code.jsp?code=SC.CH.3.5) | Apply gas laws to relationships between pressure, volume, and temperature of any amount of an ideal gas or any mixture of ideal gases using PV = nRT |

**Objective:**

**Procedure:**

1. Pour enough water into 1 flask to just cover its bottom.
2. Use the hot plate to heat the flask until the water boils and the flask fills with steam.
3. Allow the steam to escape for a short time but do NOT allow the flask to boil dry.
4. Stretch the balloon several times. Hold the mouth of the balloon open and place a small amount of liquid soap into the opening. Don’t get any on the actual opening.
5. Remove the flask from the hot plate. Hold firmly.
6. Stretch opening of balloon over lip of flask.
7. Leave flask off the hot plate and allow to cool to room temp.
8. Students record observations as it cools.
9. Make a claim as to why what occurred, occurred.
10. Draw it at the particle level.